Caught! Crystal trapping of a side-on peroxo bound to Cr(IV)

Poulsen de Sousa, David; O. Bigelow, Jennifer; Sundberg, Jonas; Que Jr., Lawrence; McKenzie, Christine Joy

Published in:
Chemical Communications

DOI:
10.1039/C4CC08785B

Publication date:
2015

Document version:
Final published version

Document license:
CC BY

Citation for published version (APA):

Go to publication entry in University of Southern Denmark's Research Portal

Terms of use
This work is brought to you by the University of Southern Denmark. Unless otherwise specified it has been shared according to the terms for self-archiving. If no other license is stated, these terms apply:

- You may download this work for personal use only.
- You may not further distribute the material or use it for any profit-making activity or commercial gain
- You may freely distribute the URL identifying this open access version

If you believe that this document breaches copyright please contact us providing details and we will investigate your claim. Please direct all enquiries to puresupport@bib.sdu.dk

Download date: 14. Jun. 2022
Caught! Crystal trapping of a side-on peroxo bound to Cr(IV)$^+$

David P. de Sousa,$^a$ Jennifer O. Bigelow,$^b$ Jonas Sundberg,$^a$ Lawrence Que Jr.$^b$ and Christine J. McKenzie$^*$$^a$

A Cr(ν)$^2$–peroxo complex crystallizes from 33% aqueous H$_2$O$_2$. The complex is a likely intermediate in catalytic disproportionation of H$_2$O$_2$ proposed to occur through a single metal site mechanism in solution – and solid state.

The study of metal–dioxygen, peroxo and oxo adducts has attracted considerable attention as such metal-activated oxygen species have emerged as the common precursors and active oxidants in a wide range of biological and non-biological oxidation systems.$^{1–5}$ The majority of first row transition metals are represented as co-factors in enzymes and proteins. However, in spite of its rich redox chemistry, reasonable availability and oxophilicity, chromium is by large absent in the natural world. Instead, it is generally considered a cell toxin capable of causing damage to DNA and RNA, through both oxidative mechanisms and adduct formation.$^6,7$ Several molecular Cr(ν)$^-$–oxo complexes have been reported in the literature over the last 40 years,$^{8–13}$ whereas only a handful Cr–dioxygen complexes have been identified. In all cases these compounds are derived from the reaction of Cr(II) precursors with O$_2$. An ion proposed to be a Cr(III)$^+$–superoxo complex$^{11}$ (bipy = 2,2′-bipyridine) was obtained by reacting coordinatively unsaturated [Cr(bipy)$_2$]$^{2+}$ in the gas phase.$^{14,15}$ Subsequently Theopold et al.$^{16}$ structurally characterized a “side-on” Cr(III)$^+$–superoxo adduct generated from the reaction of [Cr(tpena)OH]$^+$ (tpena = hydrottris[3-tert-buty1-5-methylpyrazolyl]borate) with O$_2$ in diethylether. Recently Nam et al. have reported both an “end-on” Cr(III)$^+$–superoxo$^{17}$ and a “side-on” Cr(ν)$^-$–peroxo$^{17}$ adduct supported by macrocyclic tetramethylcyclams of different ring sizes.

Molecular Cr–dioxygen species are relevant with respect to the biotoxicity of chromium, and are useful as structural mimics of more elusive metal–dioxygen adducts of the later transition metals. We have previously used amino-acid derived pyridine ligands, such as penaa$^{−}(N,N,N′$-tris[2-pyridylmethyl]ethylenediamine-N′-acetate), to generate reactive iron(ν)$^-$–oxo,$^{18,19}$ manganese(ν)$^-$–oxo complexes and putative O$_2$ adducts.$^{20,21}$ and a metal-oxidant (iodosylbenzene) adduct.$^{22}$ The instability of such high-valent iron and manganese complexes makes their structural characterization especially difficult. The isolation of more stable structural analogues based on earlier first row transition metals is a potentially fruitful strategy for gleaning structural information on such intermediates.$^{18}$ Here we describe the structural trapping of a catalytically competent Cr(ν)$^-$–peroxo adduct formed from the reaction of a Cr(III)$^+$ precursor with H$_2$O$_2$.

[Cr(III)(pena)]$^{2+}$ (1) is prepared in aqueous solution from the reaction of chromium(III)$^+$–nitrato with the sodium salt of the ligand, followed by crystallization as a diiperchlorate salt. The X-ray crystal structure shows that penaa$^{−}$ coordinates through all six donor atoms. The Cr(III)$^+$ atom displays a particularly irregular octahedral geometry, Fig. S1 (ESI†), most notable being an extremely obtuse N1–Cr–N5 angle of 115.40(4). When 1 is dissolved, UV-Vis spectra (pH 4–8) show that aqueous solutions contain an equilibrium between 1, its “pseudo hydrate”, [Cr(III)(penaH)OH]$^{2+}$ (2) and the blue congrer base [Cr(III)(pena)OH]$^-$ (3), Scheme 1. Complex 3 is the base peak in ESI-MS spectra and this speciation is supported by cyclic voltammetry, vide infra. Addition of base drives the equilibrium towards 3. When 100 eq. of H$_2$O$_2$ are added to solutions of 1, a colour change from red to violet occurs over 30 min accompanied by the evolution of dioxygen$^{23}$ consistent with modest catalase activity. The solutions are EPR silent.

This observation, together with UV-Vis spectroscopy ($λ$, 539 nm; ε, 150 M$^{-1}$ cm$^{-1}$), Fig. 1, and electrospray ionization (ESI) mass spectra – which show the dominant presence of a Cr(pena)–dioxygen adduct [CrO$_2$(pena)]$^+$ at $m/z$ 474.1 (calcd 474.1),
suggests either of the isomeric species: a triplet Cr(III)–superoxo system consisting of a Cr(III) center anti-ferromagnetically coupled to a \( \text{O}_2^- \) radical, or a triplet Cr(IV)–peroxo system. A Cr(III)–hydroperoxo adduct was ruled out, since this formulation would give rise to a simple 3/2 system presumably with a spectrum similar to that obtained for 1, Fig. S5 (ESI†). The presence of a band at 878 cm\(^{-1}\) in the resonance Raman spectrum of the purple solutions, Fig. 2(a) and Fig. S8(b) (ESI†), (full spectrum) identifies the species unambiguously as a peroxo, rather than a superxo complex.\(^{24}\) Consistently the IR spectrum of an EPR silent solid precipitated with diethyl ether–dioxane shows a strong sharp absorption at 871 cm\(^{-1}\) not seen in IR spectrum of 1, Fig. 2(b). The differences in the ligand vibrations at 1256, 1373 and 1669 cm\(^{-1}\) suggest that a pyridine arm of the ligand has become uncoordinated and protonated as seen in the salts of \([\text{V}IV\text{O}(\text{tpenaH})]^2+\) and \([\text{Fe}III_2\text{O}(\text{tpenaH})]_2^+\),\(^{18,22}\) and the formulation \([\text{Cr}IV\text{O}_2(\text{tpenaH})]^2+\) (4) could be proposed.

Single crystals of 4(ClO\(_4\))\(_2\)(H\(_2\)O\(_3\))(H\(_2\)O) were obtained only on a couple of occasions by placing the largest crystals of 1(ClO\(_4\))\(_2\)(C\(_6\)H\(_8\)O\(_2\))\(_{0.5}\) available into aqueous 33% H\(_2\)O\(_2\). Over the course of a day at 4 °C the starting material was replaced by plates of the purple peroxo complex. Bubbles of O\(_2\) were observed to slowly emerge from both the solution and the surface of the crystals. The X-ray crystal structure, Fig. 3(a), shows that the peroxide is coordinated to a seven-coordinate Cr-centre in a side-on fashion, and charge balance dictates that the cation is a Cr(IV) species. As anticipated one methylpyridyl arm is uncoordinated and protonated. The distance between
the O atoms of the O₂ moiety is 1.383(8) Å. The average Cr–L distances for 1 and 4 respectively are 2.040(1) Å and 2.014(6) Å supporting a higher oxidation for the chromium ion in 4. The +4 oxidation state is confirmed by a Bond Valence Sum (BVS) analysis that yields a BVS = 4.001. A hydrogen-bonding motif is present in the crystal, with the cations linked together by the protonated dangling pyridyl arm of one molecule and the non-coordinated carbonyl oxygen of a neighbour (PyH⁺...OOC, 1.904 Å), forming herringbone chains. The hydrogen-bonded chains stack regularly to form sheets parallel to the b-axis, forming the basis for a network of pores running along the c-axis. The pores are occupied by three co-crystallized H₂O₂ molecules per cation. The perchlorate ions H-bond predominantly to the substrate guest H₂O₂. The unusually large number of H₂O₂ molecules in the crystal lattice, and their disorder, suggest that these might be rather mobile guests. Thus we conclude that a “crystal trapping” of the reactive 4 is made possible because it is surrounded by substrate that is taken up in the solid state, selectively, from solution. A calculation of the pore dimensions after in silico removal of the H₂O₂ molecules reveals about 16% pseudo void space. These voids lie predominantly between internal lamellar surfaces with the coordinated peroxo ligands lining the surface of these sheets, Fig. 3(b). With such a spatial arrangement, it seems reasonable to assume that guest H₂O₂ molecules can creep between the sheets, while product O₂ is able to leave the crystals the same way by diffusion.

Deprotonated 4 (m/z 474.1, 4H⁻) is present in ESI mass spectra obtained on both working solutions containing excess H₂O₂, and in the spectra of isolated 4(ClO₄)₂(H₂O₂)₄ redissolved in acetonitrile. These spectra furnish some hints about the mechanism of the catalytic H₂O₂ disproportionation, and the decomposition of 4 in the absence of substrate H₂O₂, respectively. The dominant ion in both spectra is [CrIV(tpena)]⁺ (5; m/z 442.1, calc 442.1). While 5 is accessible electrochemically (vide infra), it was present, only occasionally, as a minor ion in the spectra of 1 (complex 3 dominates the spectra of 1, Fig. S2, ESI†). These observations suggest that 5 is not predominantly formed via an ionization induced one-electron reduction of 1. Collision Induced Dissociation (CID) experiments on the 474.1 ion (4H⁻) result in the loss of the mass equivalent of two oxygen atoms to directly generate 5. That no stepwise O atom loss occurs, strongly indicates that the O-O bond remains intact during this process24 and that this is the dominant route for generation of the observed Cr(II) species 5. Ions assigned to [CrIVO(tpena)]⁺ (m/z 458.1, calc 458.1) and [CrIVO(tpena)]²⁺ (m/z 229.1, calc 229.1), and its ion pair, [CrVO(tpena)]ClO₄⁺ (m/z 557.1 calc 557.1) and other high-valent O containing species appear only in the spectra of isolated 4(ClO₄)₂. These species might result from the inter- and intra-molecular O atom transfer from 4 to an unknown substrate or the ligand.25 It is noteworthy that a greater number of ligand decomposition products are observed in Fig. 4(b). On the basis of the electronic and geometrical structure of 4, the gas phase experiments through which various Cr(II), Cr(IV)-peroxo, Cr(IV)-oxo, and Cr(v)-oxo species are observed, we propose in Scheme 1, a catalytic acid–base type cycle for the relatively slow disproportionation of H₂O₂. The cycle is similar to mechanisms proposed for heme catalases.27 Indeed the possible involvement of a second coordination sphere basic pyridine to aid proton transfers is reminiscent of the role of the proximal histidine in peroxidases. The species

---

**Fig. 3** (a) ORTEP plot of the cation in 4(ClO₄)₂(H₂O₂)₄ showing 50% probability ellipsoids. (b) Ball-and-stick illustration of the crystal lattice defined by the chains of H-bonded 4 with the peroxide ligand as space filled spheres after in silico removal of H₂O₂, H₂O and ClO₄⁻. Viewed down the b-axis.

**Fig. 4** (a) ESI mass spectrum of solution of 4 generated in situ from the reaction of 1 with 200 eq. H₂O₂ in H₂O. (b) ESI mass spectrum of 4(ClO₄)₂ dissolved in CH₃CN. Important assignments not in the text: m/z 490.1 [CrIV(O)(O₂)(tpena)]⁺ (or [CrIV(O₂)(tpenaO)]⁺), 566.1 ([CrIV(tpenaO-CH₃C₆H₄N)]⁺), 350.1 ([CrIV(tpena-CH₂C₅H₄N)]⁺).
labelled by a number have been detected, and a structural analogue for 6, [V\(^{3+}\)O(tpenaH)]\(^+\) is known.\(^{18}\) The direct observation of the O\(_2\) release process step [vi] by CID suggests that the pro-catalyst 1 is not directly involved in the catalytic cycle, but requires reductive activation, by direct (i) + (ii) or indirect (iii) involvement by H\(_2\)O\(_2\) before it can enter the cycle. Time-resolved UV-Vis spectroscopy indicates that the direct (i) + (ii) mechanism is the dominant activation pathway, since the time trace for the appearance of the \(\lambda_{\text{max}}\) at 553 nm due to the chromophore of 4 can be modelled satisfactorily with a 2-step mechanism (see ESI† Fig. S7). The cyclic voltammogram of 1(ClO\(_4\))\(_2\) reveals the presence of 1 and 3 in neutral aqueous solutions, with reversible Cr\(^{2+/3+}\) redox couples at \(E_{1/2} = -885\) mV and \(-1099\) mV (vs. Ag/Ag\(^+\)), Fig. S3 (ESI†), supporting the plausibility of the involvement of 5 in the catalytic cycle. The Cr(\(\nu\))–peroxo species, 6, formed by a O atom transfer from H\(_2\)O\(_2\) to 5 (iv) is expected to be a strong base which can deprotonate H\(_2\)O\(_2\) and assist the substitution step (v),\(^{28}\) perhaps with involvement of the dangling pyridine.

In summary, we have isolated a rare example of a side-on Cr(\(\nu\))–peroxo adduct. This species is a catalytically competent species in solutions, and within crystals, for catalytic H\(_2\)O\(_2\) dismutation. Its precursor and H\(_2\)O\(_2\). This is an interesting contrast to the syntheses of reversible Cr\(^{2+/3+}\)redox couples at \(E_{1/2} = -885\) mV and \(-1099\) mV (vs. Ag/Ag\(^+\)), Fig. S3 (ESI†), supporting the plausibility of the involvement of 5 in the catalytic cycle. The Cr(\(\nu\))–oxo species, 6, formed by a O atom transfer from H\(_2\)O\(_2\) to 5 (iv) is expected to be a strong base which can deprotonate H\(_2\)O\(_2\) and assist the substitution step (v),\(^{28}\) perhaps with involvement of the dangling pyridine.

This work was supported by the Danish Council for Independent Research | Natural Sciences (grant 12-124985 to CMcK), and the US National Science Foundation (grants CHE-1058248 and CHE-1361773 to LQ). We thank Dr Mads S. Vad for recording EPR spectra and Simon Svane for performing CID experiments. Its COST CM1003 action is acknowledged for travel funding.

Notes and references

23. The gas was identified using gas chromatography. TCD detector. A compilation of stretching vibrations for Cr–O\(_2\) complexes, see Table S1 in the ESI†.
24. For details and validity of the BVS analysis, see the ESI†. A BVS analysis of the BVS analysis, see the ESI†.